



Standard Practice for Use of Mattsson's Solution of pH 7.2 to Evaluate the Stress-Corrosion Cracking Susceptibility of Copper-Zinc Alloys¹

This standard is issued under the fixed designation G37; the number immediately following the designation indicates the year of original adoption or, in the case of revision, the year of last revision. A number in parentheses indicates the year of last reapproval. A superscript epsilon (ϵ) indicates an editorial change since the last revision or reapproval.

1. Scope

1.1 This practice covers the preparation and use of Mattsson's solution of pH 7.2 as an accelerated stress-corrosion cracking test environment for brasses (copper-zinc base alloys). The variables (to the extent that these are known at present) that require control are described together with possible means for controlling and standardizing these variables.

1.2 This practice is recommended only for brasses (copper-zinc base alloys). The use of this test environment is not recommended for other copper alloys since the results may be erroneous, providing completely misleading rankings. This is particularly true of alloys containing aluminum or nickel as deliberate alloying additions.

1.3 This practice is intended primarily where the test objective is to determine the relative stress-corrosion cracking susceptibility of different brasses under the same or different stress conditions or to determine the *absolute* degree of stress corrosion cracking susceptibility, if any, of a particular brass or brass component under one or more specific stress conditions. Other legitimate test objectives for which this test solution may be used do, of course, exist. The tensile stresses present may be known or unknown, applied or residual. The practice may be applied to wrought brass products or components, brass castings, brass weldments, and so forth, and to all brasses. Strict environmental test conditions are stipulated for maximum assurance that apparent variations in stress-corrosion susceptibility are attributable to real variations in the material being tested or in the tensile stress level and not to environmental variations.

1.4 This practice relates solely to the preparation and control of the test environment. No attempt is made to recommend surface preparation or finish, or both, as this may vary with the test objectives. Similarly, no attempt is made to recommend particular stress-corrosion test specimen configurations or methods of applying the stress. Test specimen configurations that may be used are referenced in Practice G30 and STP 425.²

1.5 The values stated in SI units are to be regarded as standard. No other units of measurement are included in this standard.

1.6 *This standard does not purport to address all of the safety concerns, if any, associated with its use. It is the responsibility of the user of this standard to establish appropriate safety and health practices and determine the applicability of regulatory limitations prior to use.* (For more specific safety hazard statements see Section 8.)

2.1 *ASTM Standards:*³
D1193 Specification for Reagent Water
G30 Practice for Making and Using U-Bend Stress-Corrosion Test Specimens

2. Referenced Documents

3. Summary of Practice

4. Significance and Use

3.1 The practice consists of completely and continuously immersing a stressed test specimen in an aqueous solution containing 0.05 g-atom/L of Cu^{++} and 1 g-mol/L of NH_4^+ and of pH 7.2. The copper is added as $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ and the NH_4^+ as a mixture of NH_4OH and $(\text{NH}_4)_2\text{SO}_4$. The ratio of these latter two compounds is adjusted to give the desired pH. Exposure time, criterion of failure, and so forth, are variable and not specifically recommended.

4.1 This test environment is believed to give an accelerated ranking of the relative or *absolute* degree of stress-corrosion cracking susceptibility for different brasses. It has been found to correlate well with the corresponding service ranking in environments that cause stress-corrosion cracking which is thought to be due to the combined presence of traces of

¹ This practice is under the jurisdiction of ASTM Committee G01 on Corrosion of Metals and is the direct responsibility of Subcommittee G01.06 on Environmentally Assisted Cracking.

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² STP425, *Stress Corrosion Testing*, ASTM International.

³ For referenced ASTM standards, visit the ASTM website, www.astm.org, or contact ASTM Customer Service at service@astm.org. For *Annual Book of ASTM Standards* volume information, refer to the standard's Document Summary page on the ASTM website.

moisture and ammonia vapor. The extent to which the accelerated ranking correlates with the ranking obtained after long-term exposure to environments containing corrosives other than ammonia is not at present known. Examples of such environments may be severe marine atmospheres (Cl^-), severe industrial atmospheres (predominantly SO_2), and super-heated ammonia-free steam.

4.2 It is not possible at present to specify any particular time to failure (defined on the basis of any particular failure criteria) in pH 7.2 Mattsson's solution that corresponds to a distinction between acceptable and unacceptable stress-corrosion behavior in brass alloys. Such particular correlations must be determined individually.

4.3 Mattsson's solution of pH 7.2 may also cause stress independent general and intergranular corrosion of brasses to some extent. This leads to the possibility of confusing stress-corrosion failures with mechanical failures induced by corrosion-reduced net cross sections. This danger is particularly great with small cross section specimens, high applied stress levels, long exposure periods and stress-corrosion resistant alloys. Careful metallographic examination is recommended for correct diagnosis of the cause of failure. Alternatively, unstressed control specimens may be exposed to evaluate the extent to which stress independent corrosion degrades mechanical properties.

5. Apparatus

5.1 Any suitable inert container may be used to hold the test solution and test specimens during exposure. Glass or plastic is highly recommended. The container should be fitted with a removable top to reduce evaporation during test, thus preventing dust and other particulate matter from entering the environment, and facilitating periodic inspection of the specimens. The top should not, however, be airtight, but instead should permit reasonably free access of laboratory air to the surface of the environment. The container and top should not be transparent and should be fully opaque to light.

5.2 *Specimen Holders*, should be designed to insulate electrically the test specimens from each other and from any other bare metal. They should also be completely inert to the test environment so that leaching of soluble components or direct reaction with the test environment will not interfere with the test. The shape and form of the specimen holders and supports should be such that they do not interfere with free access to the test environment to the test specimen and its stressed surfaces.

6. Reagents and Materials

6.1 Reagent grade copper sulfate crystals ($\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$), ammonium sulfate crystals ($(\text{NH}_4)_2\text{SO}_4$), and concentrated ammonium hydroxide solution (NH_4OH) (28 to 30 % NH_3 assayed and 0.90 sp gr) conforming to the specifications of the

Committee on Analytical Reagents of the American Chemical Society shall be used.⁴

6.2 *Purity of Water*—Reagent water Type IV (Specification **D1193**) shall be used to prepare the test solution.

7. Test Solution

7.1 The concentration of the test solution shall be 0.05 g-atom/L with respect to Cu^{++} and 1.0 g-mol/L with respect to NH_4^+ . The pH of the test solution shall be $7.2 + 0.3, -0.1$ pH.

7.2 The test solution shall be prepared by completely dissolving 590.0 ± 1.0 g of $(\text{NH}_4)_2\text{SO}_4$ in 4 L of water and by completely dissolving 125.0 ± 0.5 g of $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ in 1 L of water. These two solutions should then be thoroughly mixed and 71.0 ± 0.2 mL of NH_4OH solution added, preferably with a buret. Finally, the whole should be diluted to 10.0 ± 0.1 L and allowed to age for 48 to 96 h in the test container prior to use. It is not recommended that the solution be stored for extended periods or used without the specified aging. Smaller or larger volumes of solution can be prepared using lesser amounts of reagents in the same proportions.

7.3 After aging, the pH of the test solution should be measured. If outside the range specified above, the pH may be adjusted to within the range 7.1 to 7.5 by the addition of fresh pH 4 or pH 10 Mattsson's solution. Addition of NH_4OH or H_2SO_4 to adjust pH is not recommended since the concentration of the various soluble ions can be considerably altered.

7.4 Temperature control of the test solution is not recommended. Instead, the room air temperature should be controlled at $21 \pm 3^\circ\text{C}$ and the test solution allowed to reach its equilibrium temperature with the air. No room air relative humidity control is recommended and temperature recording is not mandatory.

7.5 The volume of test solution used should be sufficiently large to avoid any appreciable change in corrosivity due to either exhaustion of corrosive constituents or accumulation of corrosion products. An arbitrary minimum ratio between the volume of test solution and the exposed specimen surface area of 3 L/dm^2 (30 mL/cm^2) is recommended.

7.6 The test solution should be changed after every 7 days of specimen exposure time by transferring the test specimens to a new batch of 48-h aged test solution. Evaporation losses should be replenished with reagent water Type IV (Specification **D1193**) as frequently as deemed necessary. As far as possible, the air above the solution should be static to minimize water and ammonia evaporation. Forced ventilation should be

⁴ *Reagent Chemicals, American Chemical Society Specifications*, American Chemical Society, Washington, DC. For suggestions on the testing of reagents not listed by the American Chemical Society, see *Analar Standards for Laboratory Chemicals*, BDH Ltd., Poole, Dorset, U.K., and the *United States Pharmacopeia and National Formulary*, U.S. Pharmacopoeial Convention, Inc. (USPC), Rockville, MD.